**Stoich / Limiting Reagent Makeup Quiz: Honors**

Given the balanced equation below, solve the following questions: (1 pt)

**\_\_\_ Na2O + \_\_\_ K3PO4 → \_\_\_ Na3PO4 + \_\_\_ K2O**

1. How many grams of sodium oxide will it take to make 45.0 grams of sodium phosphate, given an excess of potassium phosphate? (5 pt)
2. If I wanted to do this reaction and had only 34 grams of sodium oxide and 54 grams of potassium phosphate, how many grams of potassium oxide could I make? (10 pt)
3. If I performed the reaction in problem 2 and made 37.8 grams of potassium oxide, what would the percent yield for this reaction be? What would the percent yield tell you about how well I performed the lab? (4 pt)
4. Define the following terms: (2 pt each)

* limiting reagent:
* stoichiometry: